

Name _____ Period _____ Date _____ Seat _____

Unit 2 and 3 Study Guide

Objectives	Define or give an example
1. Describe the characteristics of solids, in terms of particles and their arrangement: use particle diagrams to account for motion and density differences;	
2. What is meant by “lattice”?	
3. Describe melting.	
4. Describe freezing	
5. Describe the characteristics of liquids, in terms of particles and their arrangement: use particle diagrams to account for motion and density differences;	
6. Describe evaporation.	
7. Describe condensation	
8. Describe the characteristics of gases, in terms of particles and their arrangement: use particle diagrams to account for motion and density differences;	
9. States of matter, change of state and phase change.	
10. Attractive and repulsive (repelling) forces.	
11. Explain temperature. What causes temperature?	
12. What is a fluid.	

12. <i>What is kinetic energy?</i>	
13. Relate temperature to the kinetic energy and the speed of the particles. Expansion: Contraction:	Hot: Cold:
14. Explain, at the particle level, why the level in a thermometer rises when it is placed in a warmer fluid. Explain with the three step process: heat transfer, speed of molecules, expansion/ contraction,	1. 2. 3.
15. Explain why the level in a thermometer rises when it is placed in a warmer fluid.	1. 2. 3.
16. Describe the Celsius temperature scale.	
17. State the basic tenets of the Kinetic Molecular Theory (KMT).	
18. What is heat? Explain heat flow.	
19. Explain the difference between temperature and heat.	
20. What is energy?	
21. Explain pressure. What causes pressure?	
22. Molecule	
23. Atom	