**Name: \_**

Group: 1 2

Charge: +1 +2

18 , VIII

no ions

The Periodic Table: Main Group Stable Ions



**1** Hydrogen ion



|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Group: 13 14 15 16 17  III IV V VI VII  Charge: -4 -3 -2 -1 | | | | | **2 Helium**  **He,** \_\_\_e- |
| **5 Boron**  **B, 5** e- | **6 Carbide**  **C4-** | **7** | **8** | **9** | **10 Neon** |
| **13 Aluminum ion**  **Al3+,** \_\_\_\_e- | **14** | **15** | **16, Sulfide**  **S2-, \_\_\_\_** | **17** | **18** |

H+, 0 e-

1p+,

1p+ **X**

1

**3 4**

**2**

**11 12, Magnesium**

**ion**

Mg2+, \_\_\_\_e-

3

**19 20**

1. Write the **name of the ion** and the symbol with charge (example: **S2-)**

2. Write the total number of electrons next to the ion symbol

3. Draw the energy level(s) and the correct number of electrons on each level.

4

You can use “x” or a circle for each electron.

4. Write the number of protons in the nucleus of each atom.

For Nobel Gasses (group 18, VIII ), draw atoms energy levels & fill in electrons.